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At what point did each of the buffers lose their effectiveness? ExplainHelp with ap chem lab 19: pH properties of Buffer solutions? 1 mL of 0.2 M HCl is added to 50 mL of deionized water: 0.001 litres @ 0.2mol/litre = 0.0002 moles HCl 0.0002 moles HCl / 0.051 liters = 0.0039 Molar 0.0039 Molar H+ = pH of 2.4 that's a drastic change in pH from deionized water of about 7.0 ===== when the same ...

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